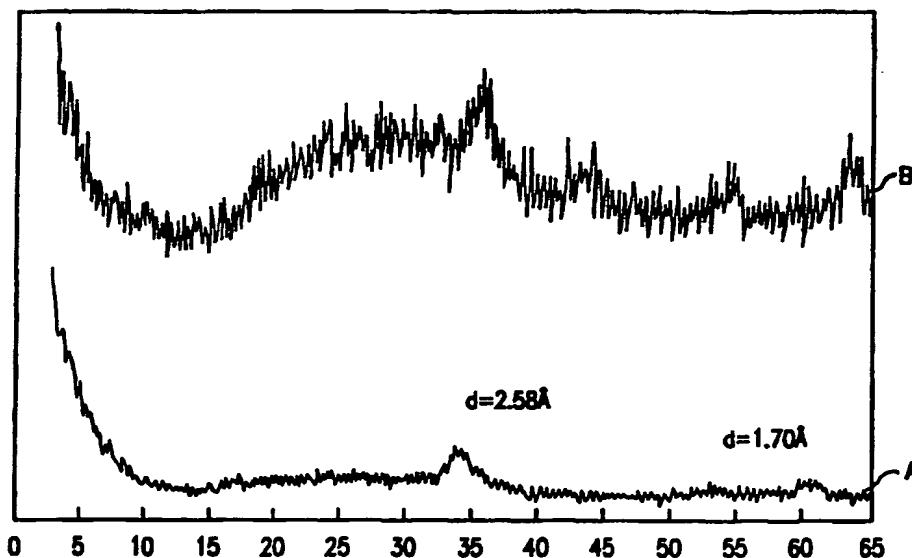


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(21) International Application Number: PCT/US98/14327 (22) International Filing Date: 10 July 1998 (10.07.98) (30) Priority Data: 08/900,389 15 July 1997 (15.07.97) US (71) Applicant: EXXON RESEARCH AND ENGINEERING COMPANY [US/US]; 180 Park Avenue, P.O. Box 390, Florham Park, NJ 07932-0390 (US). (72) Inventors: SOLED, Stuart, L.; 21 Cooks Cross Road, Pittstown, NJ 08867 (US). MISEO, Sabato; 770 County Road 579, Pittstown, NJ 08867 (US). KRYCAK, Roman; 5 Kent Court, Annandale, NJ 08801 (US). VROMAN, Hilda; 1124 Hanover Street, Piscataway, NJ 08854 (US). HO, Win-Sow, Winston; 887 Jason Court, Bridgewater, NJ 08807 (US). RILEY, Kenneth, Lloyd; 1289 Rodney Drive, Baton Rouge, LA 70808 (US). (74) Agents: PURWIN, Paul, E. et al.; Exxon Research and Engineering Company, P.O. Box 390, Florham Park, NJ 07932-0390 (US).	(81) Designated States: AL, AU, BA, BB, BG, BR, CA, CN, CU, CZ, EE, GE, HR, HU, ID, IL, IS, JP, KP, KR, LC, LK, LR, LT, LV, MG, MK, MN, MX, NO, NZ, PL, RO, SG, SI, SK, SL, TR, TT, UA, UZ, VN, YU, ARIPO patent (GH, GM, KE, LS, MW, SD, SZ, UG, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG).  Published With international search report.	

(54) Title: NICKEL MOLYBDOTUNGSTATE HYDROTREATING CATALYSTS



## (57) Abstract

A hydrodenitrogenation catalyst is prepared by decomposing a nickel (ammonium) molybdotungstate precursor and sulfiding, either pre-use or in situ, the decomposition product.

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## NICKEL MOLYBDOTUNGSTATE HYDROTREATING CATALYSTS

### FIELD OF THE INVENTION

This invention relates to new hydrodenitrogenation (HDN) catalysts. More particularly this invention relates to the decomposition product of nickel (ammonium) molybdotungstates and their use as catalysts in HDN processes.

### BACKGROUND OF THE INVENTION

As the supply of low sulfur, low nitrogen crudes decrease, refineries are processing crudes with greater sulfur and nitrogen contents at the same time that environmental regulations are mandating lower levels of these heteroatoms in products. Consequently, a need exists for increasingly efficient desulfurization and denitrogenation catalysts.

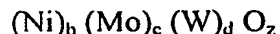
In one approach, a family of compounds, related to hydrotalcites, e.g., ammonium nickel molybdates, has been prepared. Whereas X-ray diffraction analysis has shown that hydrotalcites are composed of layered phases with positively charged sheets and exchangeable anions located in the galleries between the sheets, the related ammonium nickel molybdate phase has molybdate anions in interlayer galleries bonded to nickel oxyhydroxide sheets. See, for example, Levin, D., Soled, S. L., and Ying, J. Y., Crystal Structure of an Ammonium Nickel Molybdate prepared by Chemical Precipitation, *Inorganic Chemistry*, Vol. 35, No. 14, p. 4191-4197 (1996). The preparation of such materials also has been reported by Teichner and Astier, *Appl. Catal.* 72, 321-29 (1991); *Ann. Chim. Fr.* 12, 337-43 (1987), and *C. R. Acad. Sci.* 304 (II), #11, 563-6 (1987) and Mazzocchia, *Solid State Ionics*, 63-65 (1993) 731-35.

Now, when molybdenum is partially substituted for by tungsten, an amorphous phase is produced which upon decomposition and, preferably, sulfidation, provides enhanced hydrodenitrogenation (HDN) catalyst activity relative to the unsubstituted (Ni-Mo) phase.

### SUMMARY OF THE INVENTION

In accordance with this invention, an amorphous nickel molybdo tungstate composition is produced and exhibits enhanced hydrodenitrogenation (HDN) activity as compared to known catalyst. In essence, at least a portion but not all of the molybdenum in a nickel-molybdate system is replaced by tungsten, that is the molar ratio of molybdenum to tungsten is at least 0.01/1 and less than 0.95/1.

The composition can be further described as a bulk mixed metal oxide useful as an HDN catalyst and preferably sulfided prior to use as a catalyst, of the formula:

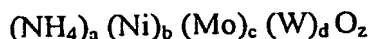


wherein the molar ratio of  $b : (c+d)$  is 0.5/1 to 3/1, preferably 0.75/1 to 1.5/1, more preferably 0.75/1 to 1.25/1;

The molar ratio of  $c:d$  is preferably  $>0.01/1$ , more preferably  $>0.1/1$ , still more preferably 1/10 to 10/1, still more preferably 1/3 to 3/1, most preferably substantially equimolar amounts of Mo and W, e.g., 2/3 to 3/2; and  $z = [2b + 6(c+d)]/2$ .

The essentially amorphous material has a unique X-ray diffraction pattern showing crystalline peaks at  $d = 2.53$  Angstroms and  $d = 1.70$  Angstroms.

The mixed metal oxide is readily produced by the decomposition of a precursor having the formula:



wherein the molar ratio of  $a:b$  is  $\leq 1.0/1$ , preferably  $0-1$ ; and  $b$ ,  $c$ , and  $d$ , are as defined above, and  $z = [a + 2b + 6(c+d)]/2$ . The precursor has similar peaks at  $d = 2.53$  and  $1.70$  Angstroms.

Decomposition of the precursor may be effected at elevated temperatures, e.g., temperatures of at least about  $300^\circ\text{C}$ , preferably about  $300-450^\circ\text{C}$ , in a suitable atmosphere, e.g., inerts such as nitrogen, argon, or steam, until decomposition is substantially complete, i.e., the ammonium is substantially completely driven off. Substantially complete decomposition can be readily established by thermogravimetric analysis (TGA), i.e., flattening of the weight change curve.

#### BRIEF DESCRIPTION OF DRAWINGS

Figure 1 is the X-ray diffraction pattern of a  $\text{NH}_4\text{-Ni-0.5Mo-0.5W-O}$  compound prepared by boiling precipitation before calcining (Curve A) and after calcining at  $400^\circ\text{C}$  (Curve B). Note that the patterns for both the precursor and the decomposition product of the precursor are quite similar with the two peaks at essentially the same place. The ordinate is relative intensity; the abscissa is two theta (degrees).

Figure 2 shows the X-ray diffraction patterns, by  $\text{CuK}\alpha$  radiation ( $\lambda=1.5405\text{\AA}$ ), of  $\text{NH}_4\text{-Ni-Mo}_{1-x}\text{-W}_x\text{-O}$  precursors wherein curve A is  $\text{Mo}_{0.9}\text{W}_{0.1}$ , curve B is  $\text{Mo}_{0.7}\text{W}_{0.3}$ , curve C is  $\text{Mo}_{0.5}\text{W}_{0.5}$ , curve D is  $\text{Mo}_{0.3}\text{W}_{0.7}$ , curve E is  $\text{Mo}_{0.1}\text{W}_{0.9}$ , and curve F is  $\text{MoO}\text{W}_1$ . The ordinate and abscissa are as described for Figure 1.

### PREFERRED EMBODIMENTS

The precursor compound can be readily prepared by one of several methods, including a variation of the boiling decomposition method used by Teichner and Astier in which a tungsten compound is added to the initial mixture of a molybdenum salt, a nickel salt and ammonium hydroxide. Direct precipitation and pH controlled precipitation may also be used to prepare the precursor compound. In all cases, however, water soluble salts of nickel, molybdenum and tungsten are employed.

Preferably, the molybdenum and tungsten salts are ammonium compounds, e.g., ammonium molybdate, ammonium metatungstate, while the nickel salt may be the nitrate or hydrated nitrates.

In the boiling decomposition method, the salts are dissolved in water to make an acidic solution, after which additional  $\text{NH}_4\text{OH}$  is added to make a basic solution. The solution is then heated to boiling to drive off ammonia and form a precipitate which is filtered and dried, e.g. at  $100\text{-}125^\circ\text{C}$ .

In the direct precipitation method, initially the molybdate and tungstate salts are dissolved in water,  $\text{NH}_4\text{OH}$  is added to form a basic solution, and the solution is warmed. A warm, e.g.,  $90^\circ\text{C}$ , nickel salt solution (aqueous) is slowly added to the initial solution, a precipitate is formed, the solution is hot

filtered and dried. In either the boiling decomposition method or the direct precipitation method, washing of the filtrate is minimized to prevent leaching.

In general, all of the components, the Ni, Mo, W,  $\text{NH}_3$ , are mixed in solution together and heated to a pH  $<7$  to form the precipitate, i.e., the precursor compound. This may be accomplished by either of two methods: (1): adding all of the components together with an excess of ammonia to dissolve the components and then heating to drive off the ammonia such that the pH  $<7$  (heating may be at less than  $100^\circ\text{C}$ , preferably about  $50-90^\circ\text{C}$ ); or (2) adding together one or more separate solutions of each component such that the final pH is  $<7$ ; in each case recovering the resulting precipitate.

In another embodiment, a binder can be added to the bulk mixed metal oxide to maintain particle integrity. The binder can be silica, alumina, silica-alumina or other materials generally known as particle binders. When utilizing a binder, the amount may range from about 1-30 wt% of the finished catalyst, preferably about 5-26 wt% of the finished catalyst.

After recovering the precursor product, regardless of preparation method, the precursor is decomposed at temperatures ranging from about  $300-450^\circ\text{C}$  in a suitably inert or air atmosphere.

The decomposed precursor can be sulfided or pre-sulfided by a variety of known methods. For example, the decomposition product can be contacted with a gas comprising  $\text{H}_2\text{S}$  and hydrogen, e.g., 10%  $\text{H}_2\text{S}/\text{H}_2$ , at elevated temperatures for a period of time sufficient to sulfide the decomposition product, usually at the point of  $\text{H}_2\text{S}$  breakthrough in the exit gas. Sulfiding can also be effected, in situ, by passing a typical feedstock containing sulfur over the decomposition product.

Any hydrocarbon containing feed which also contains nitrogen may be treated with the enhanced catalysts of this invention. Thus, the HDN process with these catalysts may range from petroleum distillates to residual stocks, either virgin or cracked, to synthetic fuels such as coal oils or shale oils. The HDN process is particularly useful with feeds containing high levels of nitrogen, e.g., at least about 500 ppm total nitrogen compounds. Nitrogen removal is at least about 50%, preferably at least about 80%.

Process conditions applicable for the use of the catalysts described herein may vary widely depending on the feedstock to be treated. Thus, as the boiling point of the feed increases, the severity of the conditions will also increase. The following table serves to illustrate typical conditions for a range of feeds.

FEED	TYPICAL BOILING RANGE °C	TEMP. °C	PRESS, BAR	SPACE VELOCITY V/V/HR	H <sub>2</sub> GAS RATE SCF/B
naphtha	25-210	100-370	10-60	0.5-10	100-2,000
diesel	170-350	200-400	15-110	0.5-4	500-6,000
heavy gas oil	325-475	260-430	15-170	0.3-2	1000-6,000
lube oil	290-550	200-450	6-210	0.2-5	100-10,000
residuum	10-50%>575	340-450	65-1100	0.1-1	2,000-10,000

While the invention described herein shows enhanced activity for hydrodenitrogenation, most HDN catalysts will also show hydrodesulfurization (HDS) activity. Consequently, the catalysts and processes described herein will be useful on feeds containing both nitrogen and sulfur, and will be particularly useful on feeds high in nitrogen.



The following examples will serve to illustrate, but not limit, this invention.

**Example 1.** Preparation of  $\text{NH}_4\text{-Ni-Mo-O}$  Phase (boiling decomposition as per Teichner and Astier procedure):

In a 1 liter flask, 26.5 g ammonium molybdate (0.15 moles Mo) and 43.6 g nickel nitrate hexahydrate (0.15 moles Ni) were dissolved in 300 cc of water so that the resulting pH equaled 4.3. To this solution, a concentrated  $\text{NH}_4\text{OH}$  solution was added. At first, a precipitate formed which on further addition of  $\text{NH}_4\text{OH}$  dissolved to give a clear blue solution with a pH of 8.3, and additional  $\text{NH}_4\text{OH}$  (~250cc) was added until a pH of 10 was reached. The solution was heated to  $90^\circ\text{C}$  for 3 h during which ammonia gas evolved and a green precipitate formed. The final pH lay between 6.8 and 7. The suspension was cooled to room temperature, filtered, washed with water and dried at  $120^\circ\text{C}$  overnight. About 18.6g of material was obtained. The sample analyzed for Ni at 26.6 wt% and Mo at 34 wt%. The X-ray diffraction spectra of the phase matches the pattern reported by Teichner.

**Example 2.** Preparation of  $\text{NH}_4\text{-Ni-Mo}_5\text{W}_5\text{-O}$  by boiling decomposition:

In a 1 liter flask, 13.2 g ammonium molybdate (0.075 moles Mo), 18.7 g ammonium metatungstate (.075 moles W) and 43.6 g nickel nitrate hexahydrate (0.15 moles Ni) were dissolved in 300cc of water so that the resulting pH equaled 4.3. To this solution, a concentrated  $\text{NH}_4\text{OH}$  solution (~600cc) was added until the pH reached 10. At this point, some precipitate remained. The solution was refluxed at  $\sim 100^\circ\text{C}$  for 3 h. During this heating, the precipitate dissolved to give a clear blue solution and on further heating, a green precipitate formed. The heating was continued until the pH reached

between 6.8 and 7. The suspension was cooled to room temperature, filtered, washed with water and dried at 120°C overnight. 18 grams of material is obtained. The X-ray diffraction spectra of the phase is given in Figure 1 showing an amorphous background with the two largest peaks at  $d=2.58$  and  $1.70\text{\AA}$ .

**Example 3.** Preparation of  $\text{NH}_4\text{-Ni-Mo}_{.5}\text{W}_{.5}\text{-O}$  by direct precipitation:

In a 1 liter flask, 17.65 g of ammonium molybdate (0.1 mole Mo) and 24.60 g of ammonium metatungstate (0.1 mole W) were dissolved in 800 cc of water giving a solution pH of  $\sim 5.2$ . To this solution 0.4 moles of  $\text{NH}_4\text{OH}$  ( $\sim 30$  cc) was added, raising the pH to  $\sim 9.8$  (solution A). This solution was warmed to 90°C. A second solution was prepared by adding 58.2 g of nickel nitrate, (0.2 moles Ni) which was dissolved in 50 cc of water (solution B) and maintained at 90°C. This solution was added dropwise at a rate of 7 cc/min into the ammonium molybdate/ammonium metatungstate solution. A precipitate begins to form after 1/4 of the solution was added. This suspension which was at a pH  $\sim 6.5$  was stirred for 30 minutes while the temperature was maintained at 90°C. The material was filtered hot, washed with hot water, and dried at 120°C. Approximately 38 g of material was recovered.

**Example 4.** Preparation of  $\text{NH}_4\text{-Ni-Mo}_{.5}\text{-Mo}_{.5}\text{W}_{.5}\text{-O}$  by controlled pH precipitation:

Two solutions were prepared with the same amounts of nickel, tungsten, molybdenum and ammonium hydroxide are described in Example 3 (solutions A and B) except that each solution contained about 700 cc of water. The two solutions were added into a separate vessel initially containing 400 cc of water held at 90°C. Solution B (the acidic solution) was pumped into the vessel at a constant rate of  $\sim 15\text{cc/min}$ , while solution A is added through a separate

pump which is under feedback PC control and set to maintain the pH at 6.5. On mixing the two solutions a precipitate forms. The slurry was stirred at 90°C for 30 minutes, filtered hot, washed with hot water, and dried at 120°C.

**Example 5.** Catalytic Evaluation Using Dibenzothiophene (DBT):

1.5-2 g of the catalysts of Examples 1-4 were placed in a quartz boat which was in turn inserted into a horizontal quartz tube and placed into a Lindberg furnace. The temperature was raised to 370°C in about one hour with N<sub>2</sub> flowing at 50 cc/m, and the flow continued for 1.5 h at 370°C. N<sub>2</sub> was switched off and 10% H<sub>2</sub>S/H<sub>2</sub> then added to the reactor at 20 cc/m, the temperature increased to 400°C, and held there for 2 hours. The heat was then shut off and the catalyst cooled in flowing H<sub>2</sub>S/H<sub>2</sub> to 70°C, at which point this flow was discontinued and N<sub>2</sub> was added. At room temperature, the quartz tube was removed and the material transferred into a N<sub>2</sub> purged glove box. Catalysts were evaluated in a 300cc modified Carberry batch reactor designed for constant hydrogen flow. The catalyst was piled and sized to 20/40 mesh and one gram was loaded into a stainless steel basket, sandwiched between a layer of mullite beads. 100 cc of liquid feed, containing 5 wt% dibenzothiophene in decalin was added to the autoclave. A hydrogen flow of 100 cc/min was passed through the reactor and the pressure was maintained at 3150kPa using a back pressure regulator. The temperature was raised to 350°C at 5-6 deg/min and run until either 50% DBT was converted or until 7 hours was reached. A small aliquot of product was removed every 30 minutes and analyzed by GC. Rate constants for the overall conversion as well as the conversion to the reaction products biphenyl (BP) and cyclohexylbenzene (CHB) were calculated as described by M. Daage and R. R. Chianelli [J. Cat. 149, 414-27 (1994)] and are shown in Table 1. As described in that article, high selectivities to cyclohexylbenzene relative to BP during the desulfurization reaction are a good indication of a catalyst with high

hydrodenitrogenation activity, whereas high selectivities of BP relative to CHB indicates a catalyst with high hydrosulfurization activity.

The results show that partial substitution of tungsten for molybdenum results in catalysts that are substantially higher for DBT conversion. A standard supported Ni-Mo on  $\text{Al}_2\text{O}_3$  catalyst is also shown for comparison. The high CHB/BP ratio suggests that the catalysts are active for HDN.

Table 1. Comparison of Activity in DBT Conversion Tests With Tungsten Addition by Different Preparation Schemes

catalyst	preparation technique	example #	$K_{\text{total}}$ @ 350°C	CHB/BP @ 350°C
$\text{NH}_4\text{-Ni-Mo-O}$	boiling decomposition	1	106	10.4
$\text{NH}_4\text{-Ni-Mo}_{.5}\text{W}_{.5}\text{-O}$	boiling decomposition	2	171	10.2
$\text{NH}_4\text{-Ni-Mo}_{.5}\text{W}_{.5}\text{-O}$	direct precipitation	3	167	12.4
$\text{NH}_4\text{-Ni-Mo}_{.5}\text{W}_{.5}\text{-O}$	controlled pH preparation	4	181	12.0
$\text{Ni}_2\text{Mo/Al}_2\text{O}_3$	impregnation		129	6.4

#### Example 6.

A series of catalysts were prepared in accordance with the general preparation scheme of example 2 (i.e., boiling decomposition) but varying the Mo and W relative ratios by changing the amount of ammonium molybdate and ammonium metatungstate added to the solutions. Decomposition was effected as described in Example 5. The catalysts so prepared are shown in Table 2 along with their catalytic activities for DBT measured as described in Example 5.

Table 2. Comparison of Activity in DBT Conversion Tests with Variation in  
Relative W and Mo content

Catalyst	Sample	ammonium molybdate (g)	ammonium metatungstate (g)	nickel nitrate hexahydrate (g)	$K_{\text{total}}$ @ 350°C	CHB/BP @ 350°C
$\text{NH}_4\text{-NiW-O}$	18983-97	0	36.95	43.62	128	11.3
$\text{NH}_4\text{-NiMo}_{.1}\text{W}_{.9}\text{-O}$	18983-125	2.65	33.62	43.62	132	14.1
$\text{NH}_4\text{-NiMo}_{.3}\text{W}_{.7}\text{-O}$	18983-101	7.94	25.87	43.62	154	11.6
$\text{NH}_4\text{-NiMo}_{.5}\text{W}_{.5}\text{-O}$	18357-109	13.17	18.74	43.62	171	10.2
$\text{NH}_4\text{-NiMo}_{.7}\text{W}_{.3}\text{-O}$	18983-95	18.54	11.09	43.62	158	11.5
$\text{NH}_4\text{-NiMo}_{.9}\text{W}_{.1}\text{-O}$	18983-92	23.83	3.69	43.62	141	10.5

The data show that the most active catalyst contains an approximately equimolar mixture of tungsten and molybdenum.

**Example 7.**

A series of catalysts were prepared as described in Example 3 (direct precipitation) in which equimolar mixtures of Mo and W were precipitated but the nickel content was varied. Decomposition was effected as described in Example 5. The catalysts so prepared are shown in Table 3 along with their catalytic activities for DBT measured as described in example 5.

Table 3. Variation of Nickel Content in  $\text{NH}_4\text{-Ni-Mo}_5\text{W}_5\text{-O}$  Catalysts

Catalyst	Sample	ammonium molybdate (g)	ammonium metatungstate (g)	nickel nitrate hexahydrate (g)	$K_{\text{total}}$ @ 350°C	CHB/BP @ 350°C
$\text{NH}_4\text{-Ni}_{0.75}\text{Mo}_5\text{W}_5\text{-O}$	19086-110	17.65	24.6	43.65	171	13.0
$\text{NH}_4\text{-Ni}_{1.0}\text{Mo}_5\text{W}_5\text{-O}$	19086-82	17.65	24.6	58.2	167	12.4
$\text{NH}_4\text{-Ni}_{1.25}\text{Mo}_5\text{W}_5\text{-O}$	19086-111	17.65	24.6	72.75	174	11.0
$\text{NH}_4\text{-Ni}_{1.5}\text{Mo}_5\text{W}_5\text{-O}$	19086-112	17.65	24.6	87.3	148	9.55

Catalytic performance does not change substantially with variations in Ni from 0.75 to 1.5, although K appears to go through a maximum at about 1.25 Ni.

**Example 8.** A series of catalysts were prepared in which the quantity of  $\text{NH}_4\text{OH}$  used in the preparation was varied. The catalysts were prepared in accordance to the procedure described in Example 3 except that the amount of  $\text{NH}_4\text{OH}$  in solution A was varied to change to  $\text{NH}_4\text{OH}/\text{Ni}$  molar ratio when the two solutions were mixed. Decomposition was effected as described in Example 5. The catalysts so prepared are shown in Table 4 along with their catalytic activities for DBT measured as described in Example 5.



Table 4. Variation in  $\text{NH}_4\text{OH}$  Addition to Preparation

Catalyst $\text{NH}_4\text{OH}/\text{Ni}$ mole ratio	Sample	ammonium molybdate (g)	ammonium metatungstate (g)	nickel nitrate hexahydrate (g)	$\text{cm}^3$ conc $\text{NH}_4\text{OH}$	$K_{\text{total}}$ @ $350^\circ\text{C}$	$K_{\text{CHB/BP}}$ @ $350^\circ\text{C}$
1:2	19086-96	17.65	24.6	43.65	6.8	102	10.5
1:1	19086-97	17.65	24.6	58.2	14	137	10.4
2:1	19086-82	17.65	24.6	72.75	30	167	12.4
3:1	19086-104	17.65	24.6	87.3	41	164	11.4
4:1	19086-106	17.65	24.6	87.3	55	161	12.1

While decomposition of the precursor compound will drive off most, if not all, of the ammonium portion of the precursor, the preparation of the precursor and the catalytic utility of the decomposition product can be affected by the amount of  $\text{NH}_4\text{OH}$  employed. Thus, the effectiveness of the decomposition product as a catalyst is enhanced when the  $\text{NH}_4\text{OH}/\text{Ni}$  ratio in the preparation of the precursor compound is from about 1:1 to about 4:1, preferably about 1.5:1 to about 4:1, and more preferably about 2:1 to about 4:1. While not wishing to be bound by any particular theory or mechanism, there is some evidence the  $\text{NH}_4\text{OH}/\text{Ni}$  ratio causes the Ni-M-W-O phase to change in the decomposition product.

**Example 9.** The catalysts of examples 1 and 2 were compared against standard supported Ni-Mo catalysts for the conversion of a LSADO (low sulfur auto diesel oil feed). This feed contained 510 wppm sulfur, 50 wppm nitrogen, and 30.6% aromatics with a gravity of  $39.8^\circ$  API. The catalysts were tested at  $579^\circ\text{F}$ , 650 psig of  $\text{H}_2$ , and 1850 SCFB/B of  $\text{H}_2$ . The relative activities of the different catalysts are summarized in Table 5.

Table 5. Relative Hydrotreating Activities on LSADO feed

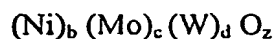
Catalyst	Relative Volumetric HDS Activity	Relative Volumetric HDN Activity
$\text{Ni}_2\text{Mo}/\text{Al}_2\text{O}_3$	1	1
$\text{NH}_4\text{-NiMo-O}$	0.25	0.50
$\text{NH}_4\text{-Ni}_{1.0}\text{Mo}_{0.5}\text{W}_{0.5}\text{-O}$	1.4	2.05

The  $\text{Ni}_2\text{Mo}/\text{Al}_2\text{O}_3$  catalyst is a standard HDN/HDS catalyst, the  $\text{NH}_4\text{-Ni-Mo}$  phase is the bulk phase with no tungsten, and the  $\text{NH}_4\text{-Ni}_{1.0}\text{Mo}_{0.5}\text{W}_{0.5}\text{-O}$  is the bulk phase with partial substitution of W for Mo. The  $\text{NH}_4\text{-NiMo-O}$  catalyst is also representative of known compounds. The catalyst of this invention is illustrated by  $\text{NH}_4\text{-Ni}_{1.0}\text{Mo}_{0.5}\text{W}_{0.5}\text{-O}$  and the data show the clear advantage of ammonium nickel tungsten molybdate for HDN activity.

## CLAIMS:

1. A composition comprising nickel molybdate in which at least a portion but less than all of molybdenum is replaced by tungsten.

2. The composition of claim 1 having the formula

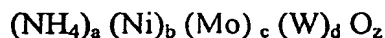


wherein the molar ratio of b: (c + d) is 0.5/1 to 3/1; the molar ratio of c:d is  $\geq 0.01/1$ ; and  $z = [2b + 6 (c+d)]/2$ .

3. The composition of claim 2 wherein the X-ray diffraction pattern is essentially amorphous with crystalline peaks at  $d = 2.53$  Angstroms and  $d = 1.70$  Angstroms.

4. The composition of claim 2 wherein the molar ratio of b: (c+d) is 0.75/1 to 1.5/1 and the molar ratio of c:d is 1/10 to 10/1.

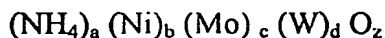
5. A mixed metal oxide precursor having the formula



wherein the molar ratio of a:b is  $\leq 1.0/1$  and b, c, and d, are as defined in claim 2, and  $z = [2b + 6 (c+d)]/2$ .

6. The decomposition product of claim 5.

7. A process for preparing a mixed metal oxide compound having the formula



wherein a, b, c, d, and z are as defined in claim 5 comprising mixing salts of Ni, Mo and W in an ammoniacal solution, and precipitating the compound by heating the solution to a pH <7.

8. The process of claim 7 wherein an ammoniacal solution containing ammonium molybdate and ammonium tungstate is mixed with a solution of a water soluble nickel salt.

9. The process of claim 7 wherein the precipitate is decomposed.

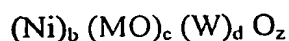
10. The process of claim 9 wherein decomposition is effected at temperatures in excess of about 300°C.

11. The product of claim 7.

12. The product of claim 9.

13. A catalytic hydrodenitrogenation process comprising contacting a hydrocarbon containing feed containing nitrogen in the presence of hydrogen and a catalyst comprising nickel molybdate in which catalyst at least a portion, but less than all, of the molybdenum is replaced by tungsten.

14. The process of claim 13 wherein the catalyst has the formula



wherein the molar ratio of b: (c + d) is 0.5/1 to 3/1; the molar ratio of c:d is  $\geq 0.01/1$ ; and  $z = [zb + 6(c + d)]/2$ .

15. The process of claim 14 wherein the X-ray diffraction pattern of the catalyst is essentially amorphous and has crystalline peaks at  $d = 2.53$  Angstroms and  $d = 1.70$  Angstroms.

16. The process of claim 14 wherein the molar ratio of b: (c + d) is 0.75/1 to 1.5/1, and the molar ratio of c:d is 1/10 to 10/1.

17. The process of claim 14 wherein the hydrocarbon containing feed contains at least about 500 ppm total nitrogen compounds.

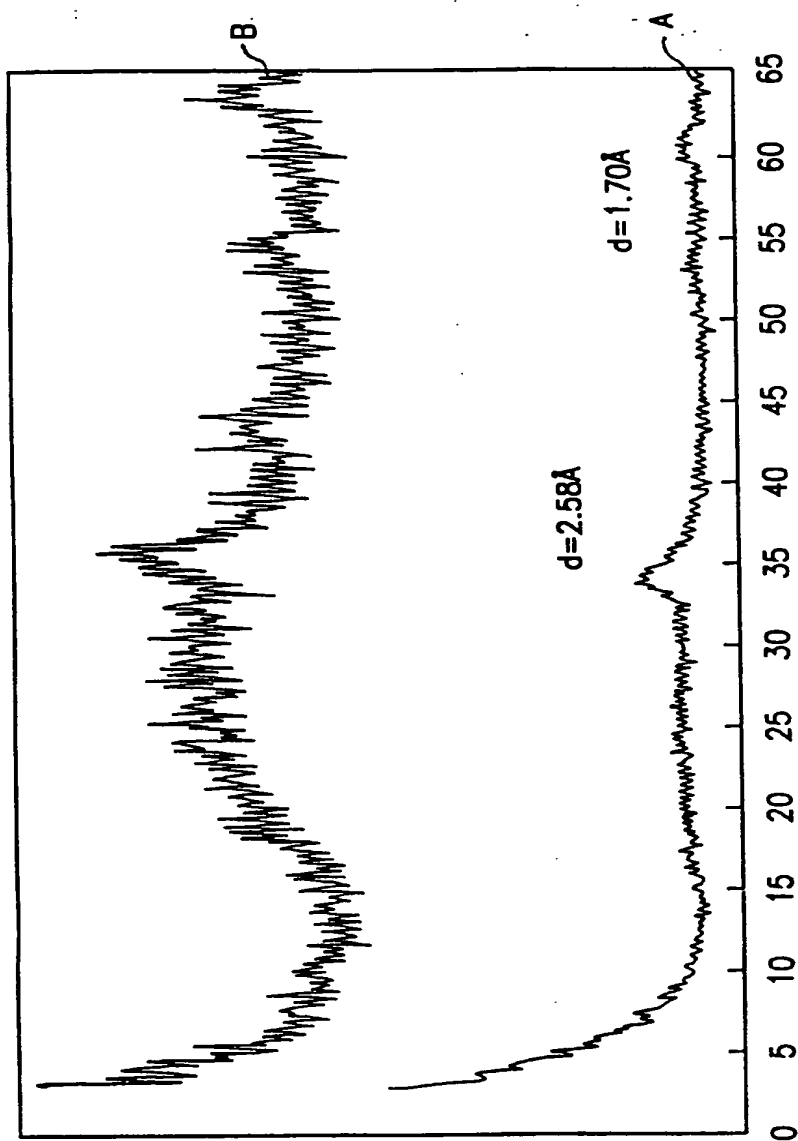


FIG. 1

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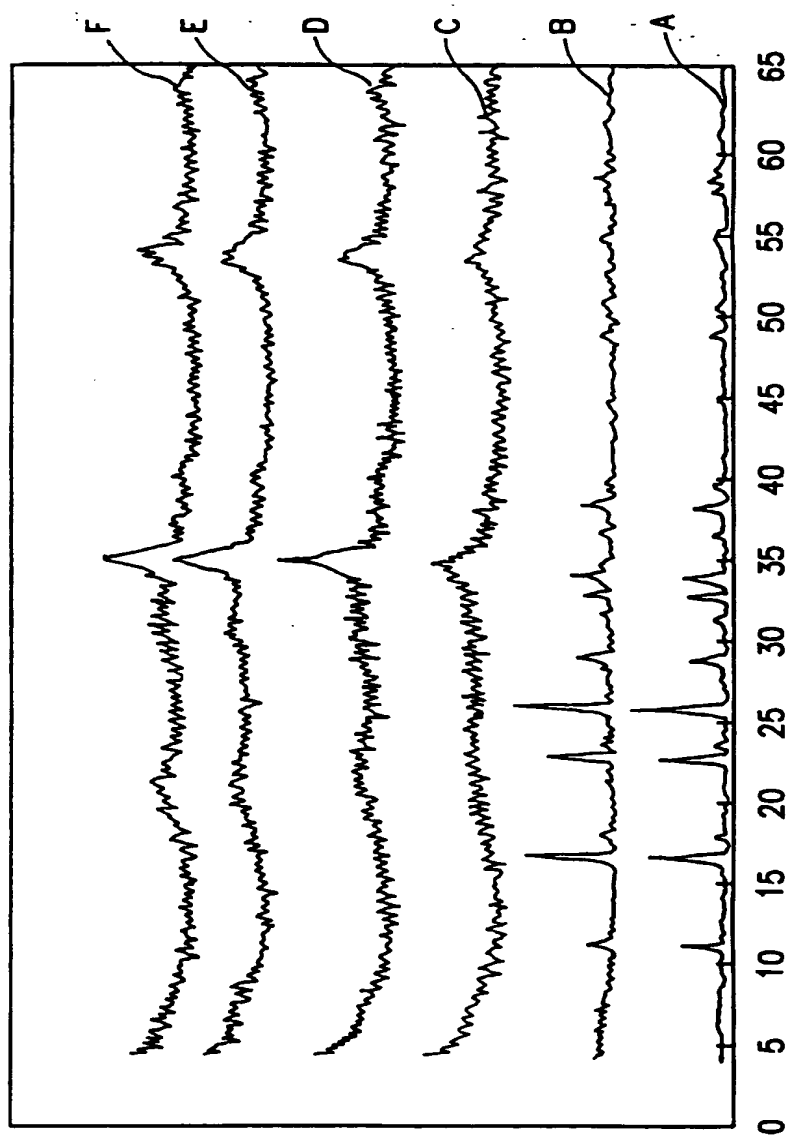


FIG. 2

# INTERNATIONAL SEARCH REPORT

International Application No

PCT/US 98/14327

## A. CLASSIFICATION OF SUBJECT MATTER

IPC 6 B01J23/883 B01J23/888 B01J37/03 C10G45/08

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 6 B01J C10G

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 3 678 124 A (STEPANOV GENNADY ARKADIEVICH ET AL) 18 July 1972 see examples 1,4 ---	1-12
X	GB 1 458 887 A (JAPAN SYNTHETIC RUBBER CO LTD) 15 December 1976 see page 6, line 11-18; example 1 ---	1-12
X	EP 0 685 260 A (NIPPON CATALYTIC CHEM IND) 6 December 1995 see example 22 ---	1
X	PATENT ABSTRACTS OF JAPAN vol. 097, no. 005, 30 May 1997 & JP 09 000929 A (SEKIYU SANGYO KASSEIKA CENTER; COSMO OIL CO LTD), 7 January 1997, see abstract --- -/-	1,13

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

### \* Special categories of cited documents:

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- "P" document published prior to the international filing date but later than the priority date claimed

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Date of the actual completion of the international search

24 September 1998

Date of mailing of the international search report

12/10/1998

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## C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	<p>M.P. ASTIER ET AL.: "Ni-Mo catalysts prepared by reduction of ammonium triammine tetranickel pentamolybdate" APPLIED CATALYSIS, vol. 72, 1991, AMSTERDAM, pages 321-29, XP002078597 cited in the application</p> <p>-----</p>	

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Information on patent family members

International Application No

PCT/US 98/14327

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		NL 7409013 A,B	07-01-1975
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		JP 8047642 A	20-02-1996
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